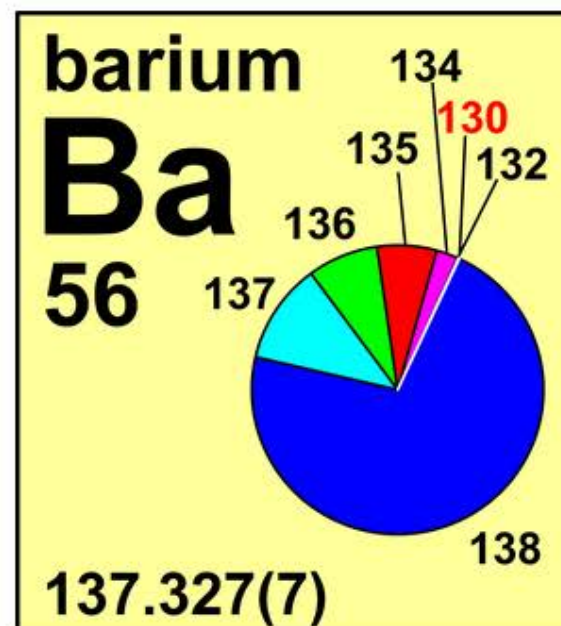


BARIUM

Isotope ↕	Atomic mass (Da) ↕	Isotopic abundance (amount fraction) ↕
^{130}Ba	129.906 32(2)	0.0011(1)
^{132}Ba	131.905 061(7)	0.0010(1)
^{134}Ba	133.904 508(2)	0.0242(15)
^{135}Ba	134.905 689(2)	0.0659(10)
^{136}Ba	135.904 576(2)	0.0785(24)
^{137}Ba	136.905 827(2)	0.1123(23)
^{138}Ba	137.905 247(2)	0.7170(29)

In 1961, the Commission recommended a value of $A_r(\text{Ba}) = 137.34$ based both on the chemical data and on mass-spectrometric abundance measurements. In 1975, the Commission reviewed the above data as well as new mass-spectrometric measurements. All of these measurements fell within a very narrow range, and recommended $A_r(\text{Ba}) = 137.33(1)$. This value was revised in 1985 to 137.327(7) as a result of new guidelines for evaluating the uncertainty of atomic weights.

SOURCE [Atomic weights of the elements: Review 2000](#) by John R de Laeter et al. *Pure Appl. Chem.* 2003 (75) 683-800
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CIAAW

Barium $A_r(\text{Ba}) = 137.327(7)$ since 1985

The name is derived from the Greek *barys* for "heavy" because it was found in the mineral heavy spar (BaSO_4). It was discovered by the Swedish pharmacist and chemist Carl Wilhelm Scheele in 1774 and first isolated by the British chemist and physicist Humphry Davy in 1808.

